

Chem I Objectives Ch 14 Condensed States

1. List the points of the kinetic molecular theory and relate each one to a property of liquids and/or solids.
2. In table form, compare the physical properties of solids and liquids to those of gases.
3. Explain what is meant by this statement, "...the state of a substance at room temperature depends on the strength of attraction between its particles."
4. Name and define three types of intermolecular forces. Rank them in order of strength.
5. Explain how viscosity and surface tension are related to intermolecular forces.
6. Relate viscosity and surface tension to temperature.
7. Describe six unusual properties of water.
8. Differentiate between crystalline and amorphous solids.
9. In table form, differentiate between metallic, molecular, ionic, and covalent network solids in terms of particles, interparticulate forces and properties. Name three examples of each and be able to recognize other examples.
10. What is meant by dynamic equilibrium?
11. Define these terms:

condensed state of matter	surface tension	viscosity	compressibility
induced dipole	vaporization	evaporation	condensation
heat of vaporization	heat of fusion	boiling point	melting point
sublimation	deposition	vapor pressure	
12. How is vapor pressure related to boiling?
13. Be able to read and interpret phase diagrams and heating curves.