

WOC Ch 4 Nomenclature Obj #5

1. Define these terms:

Ionic Bond	Covalent bond	Cation	Anion	Valence Electrons
Polyatomic ion	Monatomic ion	Oxyanions	Acids	Binary Compound
Binary ionic compound				

2. Describe the most common ion formed by each of these groups on the periodic table:

1A 2A 3A 5A 6A 7A

3. Memorize the names and formulas of the polyatomic ions on page 109.

4. Differentiate between Type I, Type II, and Type III compounds. Which ones are ionically bonded? Which are covalent compounds? Identify these compounds by type:

Magnesium acetate Phosphorous trichloride Iron (II) chloride

5. Explain how parentheses are used in formulas. Which types of compounds require parentheses?

6. Explain how Roman numerals are used in formulas. Which types of compounds require Roman numerals?

7. Explain the rules for naming acids. Give the name and formula for an example of each type of acid.

8. What does the charge on all metal ions have in common?

9. Write the name and formulas for seven Type I compounds. Include at least one that has requires parentheses. Be able to write others.

10. Write the name and formulas for seven Type II compounds. Include at least one that has requires parentheses. Be able to write others.

11. Write the name and formulas for seven Type III compounds. Use at least seven different prefixes. Be able to write others.