

Chemistry I  
WOC Ch 8 Obj # 8  
Reactions in Aqueous Solutions

1. Define these terms:

Precipitate	soluble	insoluble	dissociation	slightly soluble
Acid	strong acid	base	strong base	strong electrolyte
Salt	neutralization	spectator ion		

2. What are the driving forces that indicate a reaction is likely to occur?
3. Describe the differences between molecular equations, complete ionic equations, and net ionic equations.
4. Write a molecular equation, a complete ionic equation, and a net ionic equation for the reaction that occurs when aqueous silver nitrate and aqueous potassium chloride are combined.
5. Write a molecular equation, a complete ionic equation, and a net ionic equation for the reaction that occurs when aqueous sodium hydroxide is mixed with aqueous hydrochloric acid.
6. What are oxidation reduction reactions? Why is oxidation always paired with reduction?
7. Write an example of an oxidation reduction reaction and explain how it can be recognized as that type of reaction.
8. Write an example of a synthesis reaction and explain how it can be recognized as that type of reaction.
9. Write an example of a decomposition reaction and explain how it can be recognized as that type of reaction.
10. Write an example of a single replacement reaction and explain how it can be recognized as that type of reaction.
11. Write an example of a double replacement reaction and explain how it can be recognized as that type of reaction.
12. Write an example of a combustion reaction and explain how it can be recognized as that type of reaction.
13. Find three equations in your book that can be classified as more than one type of reaction. Copy the equations and indicate which types of reactions each is.